

Ion-ion correlation and charge reversal at titrating solid interfaces

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Confronting grand canonical titration Monte Carlo simulations (MC) with recently published titration and charge reversal (CR) experiments on silica surfaces by Dove *et al.* [1] and van der Heyden *et al.* [2], we show that ion-ion correlations quantitatively explain why divalent counterions strongly promote surface charge which, in turn, eventually causes a charge reversal (CR). Titration and CR results from simulations and experiments are in excellent agreement without any fitting parameters. This is the first unambiguous evidence that ion-ion correlations are instrumental in the creation of highly charged surfaces and responsible for their CR. Finally, we show that charge correlations result in “anomalous” charge regulation in strongly coupled conditions in qualitative disagreement with its classical treatment.

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A key parameter for understanding electrostatic interactions in charged suspensions is the surface charge density, σ [3, 4]. This parameter may control phase separation, phase transition [5], promote dissolution [6], control particle growth [7], colloidal stability [8, 9] and membrane selectivity [10] and play a significant role in catalysis and fluid transport through ionic channels [11, 12]. Typically a surface becomes charged through a titration process where surface groups ionize, *e.g.* the titration of silanol groups $-\text{Si-OH} \rightarrow -\text{Si-O}^- + \text{H}^+$.

Ion-ion correlation in solutions of charged colloids has attracted much interest due to its ability to generate net attractive interaction between equally charged colloids. This is, however, only one aspect of ion-ion correlation and an equally important aspect is its facilitation of surface titration processes. An example of vital importance is the surface titration of calcium silicate hydrates in cement paste, which reach a very high surface charge density due to ion-ion correlation among the calcium counterions. The high surface charge density is then responsible for the setting of cement [13, 14, 15].

An additional effect, due to ion-ion correlation, is when a charged surface immersed in a multivalent electrolyte more than compensates its charge resulting in a *charge reversal* (CR) [16, 17]. That is, the surface attracts counterions in excess of its own nominal charge and the apparent charge seen a few Ångström from the surface appears to be of opposite sign to the bare surface. CR is experimentally characterized by several methods, *e.g.* streaming current [2] and can induce repulsion between oppositely charged surfaces [18, 19]. Despite a vast number of studies the origin of charge reversal is still controversial as highlighted in a recent review article [20], where Lyklema points out the lack of discrimination be-

tween the *chemical* (specific chemical adsorption) and *physical* (Coulombic interactions including charge correlations) origin of CR.

The charging of interfaces (mineral or organic) has traditionally been described by the *surface complexation model* (SC) [21, 22, 23], where titration of surface groups is described by an ensemble of mass balance equilibria with associated equilibrium constants. The SC is based on the Poisson-Boltzmann (PB) equation augmented with a so-called Stern layer. The PB equation, however, neglects ion-ion correlation and *e.g.* fails to predict the attraction between equally charged surfaces [24, 25] as well as the repulsion between oppositely charged ones [19]. Despite its widespread use the limitations of the SC model in highly coupled systems have never been examined. Moreover, although multivalent ions are ubiquitous in many systems, there are only few experimental studies with titration of metal oxide particles in the presence of such ions [1, 26, 27].

Recently a grand canonical titration simulation method was developed [28] and applied to a calcium silicate hydrate (C-S-H) solution [13]. An excellent agreement was found between experiment and simulation for the titration process as well as for the CR. Unfortunately, due to the solubility range of C-S-H, the full titration curve is not accessible experimentally.

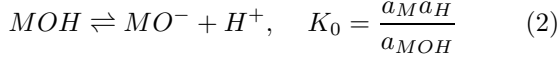
In the present letter, we scrutinize the prediction from our MC simulations for the charging process and CR by comparing to experimental data for silica particles dispersed in sodium and calcium chloride solutions [1]. In particular, we focus on the contribution of the correlations (ion-ion and ion-site). We also present a simulated charge regulation of two parallel silica surfaces in a Ca^{2+} salt solution when decreasing their separation. The experimental investigation of surface titration of silica fume particles by Dove *et al.* [1] and CR by multi-

valent counterions from a streaming current analysis in silica nanochannels by van der Heyden *et al.* [2] will serve as references.

The simulations are based on the Primitive Model with the solvent as a structureless medium characterized by its relative dielectric permittivity, ϵ_r . All charged species are treated as charged hard spheres and the interaction, between two charges i and j separated a distance r is,

$$u(r) = \frac{Z_i Z_j e^2}{4\pi\epsilon_0\epsilon_r r} \quad \text{if } r > d_{hc} \quad (1)$$

otherwise $u(r) = \infty$. Z_i is the ion valency, e the elementary charge, ϵ_0 the dielectric permittivity of vacuum and d_{hc} the hard sphere diameter of an ion always equal to 4 Å. The particles are modeled as infinite planar walls with explicit titratable sites with the intrinsic dissociation constant K_0 distributed on a square lattice. The titration process can be described as,



where the a 's denote the activities of the species. The model is solved using a Grand Canonical Monte Carlo method [29], *i.e.* at constant pH and chemical potential of the ions. In addition to ordinary MC moves the titratable sites are allowed to change their charge status. We imagine the deprotonation of a surface site as a two-step process: the release of a proton from the surface followed by an exchange of an ion pair (H^+ , B^-) with the bulk. The corresponding Boltzmann factor for the trial energy can be expressed as,

$$\exp(-\beta\Delta U) = \frac{N_B}{V} \exp[-\beta(\mu_B + \Delta U^{el}) + \ln 10(pH - pK_0)] \quad (3)$$

where μ represents the chemical potential of a particular ion, V is the volume, N_B the number of ions B^- and ΔU^{el} the change in electrostatic energy. An analogous expression holds for protonation, for more details see [13]. To establish the model parameters for a silica/solution interface, we turn our attention to the case of silica in 1:1 salt solution. The following set of parameters: $pK_0 = 7.7$, $\rho_s = 4.8/\text{nm}^2$, the surface site density, and $d_{is} = 3.5$ Å, the minimum separation between an ion and a surface site, have been found to give a perfect description of the charging process, see Fig.1. The surface charge density, σ increases, in absolute value, with pH, but is always smaller than in the ideal case due to the strong electrostatic repulsion between the deprotonated silanol groups. A change in the salt concentration modulates the screening of the electrostatic interactions which, in turn, results in a change in σ .

When the sodium counterions are replaced by divalent calcium ions, the simulations predict a strong increase in the surface charge density of silica in perfect agreement with the titration data of Dove *et al.* [1] - see

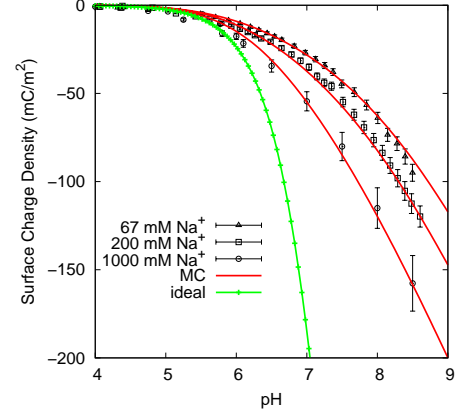


FIG. 1: Simulated (MC) and experimental (symbols) [1] surface charge density versus pH for silica particles dispersed in a sodium (1:1) salt solution at different concentrations. The ideal titration curve (ideal) is given as a reference

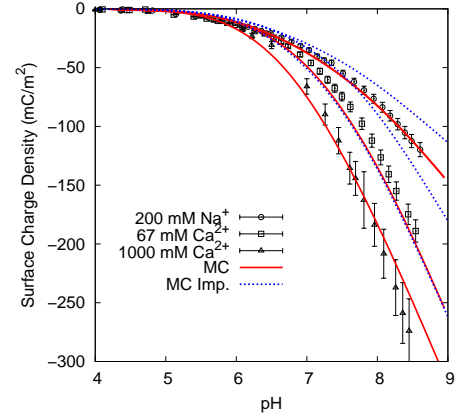


FIG. 2: Simulated (MC) and experimental (symbols) [1] surface charge density versus pH for silica particles dispersed in a calcium (2:1) and sodium (1:1) salt solution at various concentrations. The corresponding simulated titration curves for particles having implicit sites (MC imp.) are also given for comparison

Fig.2. A similar behavior has been observed from titration experiments on titania particles [27] in multivalent electrolytes and, indirectly, from proton release measurements following adsorption of multivalent ions (or polyelectrolytes) [30] emphasizing the universality of the phenomenon. The classical SC approach fails to predict the effect, unless a complexation reaction of calcium with the silanol surface groups is invoked. Here, we unambiguously show that the charge correlations (ion-ion and ion-site), without any additional fitting parameters, are enough to explain the increased silanol group ionization in the presence of Ca^{2+} ions. Thus, the complexation constants often invoked in the SC has no physical basis.

In an attempt to distinguish the effect of ion-ion from

ion-site correlations a calculation was performed using a smeared out surface site density - see Fig.2. A significant drop in $|\sigma|$ is observed, which becomes more pronounced when the counterion valence is high, although the qualitative behavior of σ upon increasing the counterion valency remains the same. Thus, the discrete nature of the surface groups have a quantitative but not qualitative effect and we can conclude that ion-ion correlations are more important than ion-site correlations, in contradiction with recent approximate studies [30, 31].

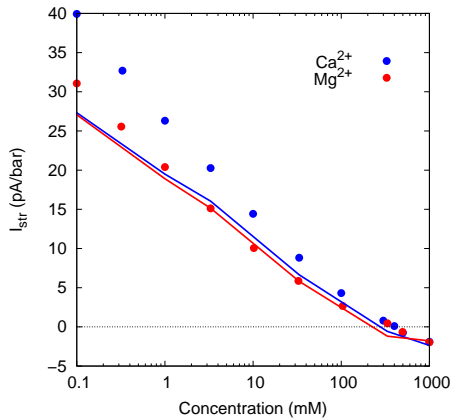


FIG. 3: Comparison between simulated (line) and experimental (points) [2] streaming currents of a silica channel as a function of electrolyte concentration of a CaCl_2 and MgCl_2 salt at pH 7.5. In the simulations, the site density is 4.8 nm^{-2} , $\text{p}K_0 = 7.7$, $d_{is}(\text{Ca}^{2+}) = 3.5 \text{ \AA}$, $d_{is}(\text{Mg}^{2+}) = 3.1 \text{ \AA}$.

We can also compare our simulation data to the streaming current, I_{str} , experiments through silica channels performed by van der Heyden *et al.* [2] in presence of di- and tri-valent counterions at pH 7.5. I_{str} is computed from the simulated local charge density, $\rho(x)$. For a channel of height h and width w with $h \gg w$, I_{str} can be defined as,

$$I_{str} = 2w \int_0^{h/2} \rho(X)v(X)dX \quad \text{with} \quad X = x - \frac{3}{2d_{hc}} \quad (4)$$

where x is the distance normal to the silica surface and $v(x)$ is the local fluid velocity, approximated as a Poiseuille flow with boundary conditions: $v(X=0) = 0$ and $v'(X=h/2) = 0$. The zero velocity is arbitrarily set at $x = 3/2d_{hc}$ where by definition we also find the electrokinetic potential. The simulations predict for both Mg^{2+} and Ca^{2+} a monotonic decrease in I_{str} upon increasing salt concentration as also found experimentally - see Fig.4. A sign reversal in the streaming potential is predicted at $c_{CR} \approx 310 \text{ mM}$ for Ca^{2+} and $c_{CR} \approx 280 \text{ mM}$ for Mg^{2+} in good agreement numbers observed experimentally [2]. What is more, the magnesium value is predicted to be below that of calcium in agreement with experiments [2] due to a higher silica charge. Simulations were also performed for trivalent

cations, where CR is predicted at sub-millimolar concentrations, $c_{CR} \approx 400 \mu\text{M}$, although times higher than what was observed with cobalt(III)sepulchrate [2].

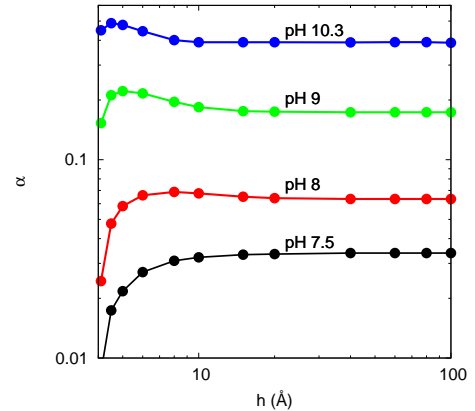


FIG. 4: The degree of ionization, α , as a function of surface separation at various pH values. The site density is 4.8 nm^{-2} and $\text{p}K_0 = 7.7$ and the surfaces are in contact with a 0.2 mM CaX_2 solution.

We now investigate the surface charging process of two parallel planar silica surfaces upon decreasing their separation, that is *the charge regulation*. The dependence of α on separation, h , in the presence of 0.2 mM Ca^{2+} at various pH is shown in Fig. 4. At low pH, a monotonic decrease of α with decreasing h , in conjunction with an increased osmotic repulsion (not shown), is found in qualitative agreement with previous mean-field studies [32, 33]. This result can be rationalized as a balance between entropy and energy. That is, when pH is well below $\text{p}K_0$, α is small and the free energy is dominated by the entropic term causing the particles to repel. As a result, the system minimizes its free energy by a reduction of the surface ionization. Conversely, atypical α vs h curves are predicted for surfaces immersed in divalent electrolytes at large pH values ($\text{pH} \geq \text{p}K_0$), where α is found to increase when shortening h . The degree of dissociation reaches a maximum at $h_{max} \approx 6 \text{ \AA}$, whereafter it decreases and annihilates at contact. The maximum in α corresponds exactly to the minimum of the potential of mean force (not shown). In this case, the energy (induced by the ion-ion correlations) dominates, which leads to an attraction between the particles and, subsequently, to an increase of the surface ionization for short h , which in turn strengthen the attraction. Close to contact the entropy regains in importance and eventually causes σ to collapse. However, when the Ca^{2+} concentration is increased to more than 10 mM (not shown), the rise in α is dramatically strengthen while the drop at close to contact is hardly measurable. In other words, the surfaces retain and even increase their charges all the way to contact.

This charge regulation behaviour may partly explain why Ca^{2+} promotes the experimentally observed dissolution of amorphous and crystalline silica surface under stress [34]. Indeed, as recently rationalized by Dove *et al.* [6], the dissolution proceeds through a nucleation process following the same mechanistic theory as developed for growth and it is shown to be promoted by surface ionization.

Ion-ion correlation also seems to be relevant for membrane selectivity. As an example, we have calculated the L-type Ca-channel selectivity at various pH in a reservoir containing always 100 mM Na^+ to which is added either 1 mM or 100 μM Ca^{2+} . The Ca-channel was modeled as an infinite slit pore with $h = 9 \text{ \AA}$, with walls decorated with titratable carboxylic groups, $\text{p}K_0 = 4.8$ and $\rho_s = 3/\text{nm}^2$. The selectivity, ξ , was calculated as the ratio between the average calcium and sodium concentrations. At pH 7 and 100 μM of Ca^{2+} we found $\xi=0.96$ and $\alpha = 0.70$, while at 1 mM Ca^{2+} the values increased to $\xi=3.7$ and $\alpha = 0.84$. Boda *et al.* [12] tried to mimic the selectivity with a constant ionization of the $-\text{COOH}$ groups, but were then forced to decrease the relative dielectric permittivity of the protein to 10 in order to fit the experimental data. When pH is reduced to 5, ξ drops to 0.10 and 0.15 at low and high calcium content, respectively in agreement with the observed loss in membrane selectivity at low pH [35].

To summarize, we have shown by confronting Monte Carlo simulations with independent experiments that the dominating interaction which controls the charging process and charge reversion of silica in multivalent electrolyte is of purely electrostatic origin and strongly dependent on ion-ion correlations. In particular, we have demonstrated that the charging process and charge reversal are intimately related. We have also demonstrated that the accepted view on charge regulation as a monotonic drop in $|\sigma|$ with decreasing separation is qualitatively wrong in highly coupled systems. Instead, $|\sigma|$ is found to increase with decreasing h . This has profound influence for the stability of colloidal particles and for the early setting of normal Portland cement.

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- [1] P. M. Dove and C. M. Craven, *Geochim. Cosmochim. Acta* **69**, 4963 (2005).
 - [2] F. H. J. van der Heyden, D. Stein, K. Besteman, S. G. Lemay, and C. Dekker, *Phys. Rev. Lett.* **96**, 224502 (2006).
 - [3] B. Honig and A. Nicholls, *Science* **268**, 1144 (1995).

- [4] K. A. Hunter and P. S. Liss, *Nature* **282**, 823 (1979).
- [5] J. Chang, P. Lesieur, M. Delsanti, L. Belloni, C. Bonnet-Gonnet, and B. Cabane, *J. Phys. Chem.* **99**, 15993 (1995).
- [6] P. M. Dove, N. Han, and J. J. D. Yoreo, *Proc. Nat. Acad. Sci.* **102**, 15357 (2005).
- [7] S. Elhadj, J. J. D. Yoreo, J. R. Hoyer, and P. M. Dove, *Proc. Nat. Acad. Sci.* **103**, 19237 (2005).
- [8] B. V. Derjaguin and L. Landau, *Acta Physicochim URSS* **14**, 633 (1941).
- [9] E. J. W. Verwey and J. T. G. Overbeek, *Theory of stability of lyophobic colloids* (Elsevier, Amsterdam, 1948).
- [10] K. Leung, S. B. Rempe, and C. D. Lorenz, *Phys. Rev. Lett.* **96**, 095504 (2006).
- [11] D. Stein, M. Kruithof, and C. Dekker, *Phys. Rev. Lett.* **93**, 035901 (2004).
- [12] D. Boda, M. Valisko, B. Eisenberg, W. Nonner, D. Henderson, and D. Gillespie, *Phys. Rev. Lett.* **98**, 168102 (2007).
- [13] C. Labbez, B. Jönsson, I. Pochard, A. Nonat, and B. Cabane, *J. Phys. Chem. B* **110**, 9219 (2006).
- [14] A. Delville, R. Pellenq, and J. Caillol, *J. Chem. Phys.* **106**, 7275 (1997).
- [15] B. Jönsson, A. Nonat, C. Labbez, B. Cabane, and H. Wenneström, *Langmuir* **21**, 9221 (2005).
- [16] R. M. Pashley, *J. Colloid Interface Sci.* **102**, 22 (1984).
- [17] Sjöström, T. Åkesson, and B. Jönsson, *Berichte der Bunsen-Gesellschaft* **100**, 889 (1996).
- [18] K. Besteman, M. A. G. Zevenbergen, H. A. Heering, and S. G. Lemay, *Phys. Rev. Lett.* **93**, 170802 (2004).
- [19] M. Trulsson, B. Jönsson, T. Åkesson, J. Forsman, and C. Labbez, *Phys. Rev. Lett.* **97**, 068302 (2006).
- [20] J. Lyklema, *Colloids Surf. A* **291**, 3 (2006).
- [21] T. Hiemstra and W. H. VanRiemsdijk, *J. Colloid Interface Sci.* **133**, 91 (1984).
- [22] D. A. Sverjensky, *Nature* **364**, 776 (1993).
- [23] M. Borkovec, *Langmuir* **13**, 2608 (1997).
- [24] O. Zohar, I. Leizeron, and U. Sivan, *Phys. Rev. Lett.* **96**, 177802 (2006).
- [25] L. Guldbrand, B. Jönsson, H. Wenneström, and P. Linse, *J. Chem. Phys.* **80**, 2221 (1984).
- [26] M. Karlsson, C. M. Craven, P. M. Dove, and W. H. Casey, *Aquatic Geochem.* **7**, 13 (2001).
- [27] M. R. Böhmer, Y. E. A. Sofi, and A. Foissy, *J. Colloid Interface Sci.* **164**, 126 (1994).
- [28] C. Labbez and B. Jönsson, *Lect. Note Comp. Sci.* **4699**, 66 (2007).
- [29] D. Frenkel and B. Smit, *Understanding molecular simulation* (Academic Press, London, 1996).
- [30] J. Pittler, W. Bu, D. Vaknin, D. J. McGillivray, and M. Lösche, *Phys. Rev. Lett.* **97**, 046102 (2006).
- [31] J. Faraudo and A. Travesset, *J. Phys. Chem. C* **111**, 987 (2007).
- [32] B. W. Ninham and V. A. Overbeek, *J. Theor. Biol.* **31**, 405 (1971).
- [33] J. G. Kirkwood and J. B. Shumaker, *Proc. Nat. Acad. Sci.* **38**, 863 (1952).
- [34] A. Anzalone, J. Boles, G. Greene, K. Young, J. Israelachvili, and N. Alcantar, *Chem. Geology* **230**, 230 (2006).
- [35] B. P. Delisle and J. Satin, *Biophys. J.* **78**, 1895 (2000).